Assigning Oxidation Numbers ➔ KEY

Determine the oxidation numbers to ALL of the elements in each of the compounds or ions below:

1) HCl
   H = +1
   Cl = −1

8) H₂O₂
   H = +1 (each)
   O = −1 (each)

15) LiH
   Li = +1
   H = −1

2) KNO₃
   K = +1
   N = +5
   O = −2 (each)

9) PbO₂
   Pb = +4
   O = −2 (each)

16) MnO₂
   Mn = +4
   O = −2 (each)

3) OH⁻
   O = −2
   H = +1

10) NaHSO₄
   Na = +1
   H = +1
   S = +6
   O = −2 (each)

17) OF₂
   O = +2
   F = −1 (each)

4) Mg₃N₂
   Mg = +2
   N = −3

11) H₂SO₃
   H = +1 (each)
   S = +4
   O = −2 (each)

18) NH₃
   N = −3
   H = +1 (each)

5) KClO₃
   K = +1
   Cl = +5
   O = −2 (each)

12) H₂SO₄
   H = +1 (each)
   S = +6
   O = −2 (each)

19) Na
   Na = 0

6) Al(NO₃)₃
   Al = +3
   N = +5 (each)
   O = −2 (each)

13) BaO₂
   Ba = +2
   O = −1 (each)

7) S₈
   S = 0

14) KMnO₄
   K = +1
   Mn = +7
   O = −2 (each)